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X Lab Data . 1. Measure Total # KJ (calorimetry), Using  $Q = C \dots$  Hess's Law . The Total Enthalpy Change For A Chemical Or Physical Change Is The Same Whether It Takes Place In One Or Several Steps. • "state Function" Version 1: A 1th, 2024

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Standard States The Thermodynamic Standard State Of A Substance Is Its Most Stable Pure Form Under Standard Pressure (1 Bar) And Some Speci C Temperature (298 K Unless Otherwise Speci Ed). I For A Pure Substance In The Liquid Or Solid Phase, The Standard State Is Th 3th, 2024

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Capacity To Do Work Or Supply Heat. Energy Has No  
Mass Or Volume. Energy Has No Mass Or Volume.  
Chemical Potential Energy Is Energy Stored In  
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$C_2H_5OH(l) + 3 O_2(g) \rightarrow 2 CO_2(g) + 3 H_2O(g)$   $\Delta H_{rxn} = -1366.8 \text{ kJ/mol}$   
A 100.0 mL Sample Of 0.300 M NaOH Is Mixed With A 100.0 mL  
Sample Of 0.305 M  $HNO_3$  In A Coffee Cup  
Calorimeter. If Both Solutions Were Initially At 35.0 °C  
And The Temperature Of The Resulting Solution Was  
Recorded As 37.0 °C, Determine The  $\Delta H_{rxn}$  (in Units Of  
kJ/mol). 4th, 2024

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