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12): Sections 5.1 (p298 - 305); P294 #3, 6, 7; P302  
#10 Answers: 1. 7.35 KJ 2.  $3.1 \times 10^2$  KJ 3. 135 J 4.  
43.5°C 5.  $C = 0.235 \text{ J/g} \cdot ^\circ\text{C}$  6.  $25.91 \text{ J} \cdot \text{Mol}^{-1} \cdot ^\circ\text{C}^{-1}$  7.  
A)  $3.8 \times 10^3 \text{ J}$  B) 53 KJ/mol 8. 23.41°C. Title: SCH4U-  
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Measuring And Expressing Enthalpy Changes  
Calorimetry Calorimetry Is The Precise Measurement  
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Heat Is Given Off By The System Is Equally Absorbed By The Surroundings. The Same Applies When 2th, 2024

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Where  $CH_0$  Is The Heat Of Combustion Of Benzoic Acid (given As  $-6318 \text{ Cal G}$ ),  $M$  Is The Mass Of The Benzoic Acid Sample ( $0.9799 \text{ G}$ ),  $E_3$  Is The Heat Of Combustion Of The Wire (calculable From The Heat Of Combustion Of The Wire,  $-2.3 \text{ Cal Cm}$  And The Length Of The Wire,  $15.1 \text{ Cm}$ ), And  $T$  Is The Previously Calculated Temperature Rise,  $2.77252 \text{ C}$  . Combining 3th, 2024

### **Experiment 9: CALORIMETRY**

$\Delta T = \text{Change In Temperature Of The Solution} = T_{\text{Final}} - T_{\text{Initial}}$  In This Experiment Mass Of The Solution Will Be Calculated Using The Volume Of The Solution And Its Density (assumed To Be  $0.997 \text{ G ML}^{-1}$ ) Heat Of Neutralization Is The Same Value As  $Q_{\text{Soln}}$  But With The Opposite Sign. The Solution Is Absorbing Heat So  $Q_{\text{Soln}}$  Is Positive. 2th, 2024

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Of The ASTM E84 Test Method. The Work Had Been Done In 1974, But Was Not Published Until 1977. There Were Earlier Efforts By Hinkley Et Al. (1968) And By Christian And Waterman (1971), Discussed Below, In Which The Re-relationship Between Oxygen Depletion In A Hooded Ex-haust Was Directly 1th, 2024

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The Increase In Temperature Be? We Are Given Both The Number Of Moles Of Compound And The Molar Heat Of Combustion So By Multiplying Them Together We Can Find The Heat Generated.  $(0.875 \text{ Mol})(1350 \text{ KJ/mol}) = 1180 \text{ KJ} = 118000 \text{ J}$  Plug This Value Into Our Equation And Then Rearrange To Find The Change In The 3th, 2024

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1. Using The Graduated Cylinder, Obtain 50 ML Of Water And Carefully Pour It Into The Soda Can. 2. Determine The Mass Of Water And The Can. Record The Mass Of Water In The Data Table (hint: Density Of Water = 1 G/mL). 3. Hold The Paper Clip Horizontally And Bend The Outer End Upwards Until 2th, 2024

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