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$2 \rightarrow \text{CaSiO}_3$ B. $3\text{Fe}_2\text{O}_3 + \text{CO} \rightarrow 2\text{Fe}_3\text{O}_4 + \text{CO}_2$ C. $\text{FeO} + \text{SiO}_2 \rightarrow \text{FeSiO}_3$ D. $\text{FeO} \rightarrow \text{Fe} + \frac{1}{2}\text{O}_2$ A. A B. D C. C And D D. A And D Answer: C Solution: In Metallurgy Of Iron, CaO Is Used As Flux Which Is Used To Remove The Impurities Of SiO_2 , $\text{CaO} + \text{SiO}_2 \rightarrow \text{CaSiO}_3$. Also Here Fe_2O_3 Is Reduced By CO To Fe_3O_4 Which Is Further Reduced To ... 4th, 2024

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