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2.2 L Of Solution.  $\text{HCl} \rightarrow \text{H}^+ + \text{Cl}^-$  Molarity Of HCl (which Is The  $\text{H}^+$  Same As  $\text{H}_3\text{O}^+$ )  
 $= 6.2 \times 10^{-3} \text{ Mol} / 2.2 \text{ L} = 2.9 \times 10^{-3} \text{ M}$  [H 3 Apr 4th, 2024.

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[Type Text] LESSON 1 SOME BASIC CONCEPT OF CHEMISTRY Chemistry 1) 40 ml Of  
Methane Is Completely Burnt Using 80 ml Of Oxygen At Room Temperature The  
Volume Of Gas Lez. After Cooling To Room Temperature Is (a) 40 ml  $\text{CO}_2$  Gas (b) 40  
ml  $\text{CO}_2$  Gas And 80 ml  $\text{H}_2\text{O}$  Gas (c) 60 ml  $\text{CO}_2$  Gas And 60 ml  $\text{H}_2\text{O}$  Gas (d) 120 ml  
 $\text{CO}_2$  Gas Feb 4th, 2024

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